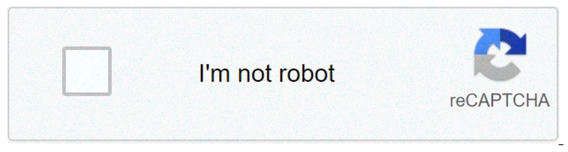

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Atomic radii of fluorine and neon

Atomic radii of fluorine and neon in pm. Atomic radii of fluorine and neon in a degree unit are respectively. Atomic radii of fluorine and neon in angstrom units are respectively given by. Atomic radii of fluorine and neon in a^(@) units are respectively. Atomic radii of fluorine atom and neon atom in angstrom units. Atomic radii of fluorine and neon in angstrom units respectively. Atomic radii of fluorine and neon in a degree are respectively given by. Atomic radii of fluorine and neon in pm are respectively given by.

Q) Log in or log in to enter a comment. Please log in or register to answer this question. 24252627282930313233343536373839404142Try our Mini Coursemaster important topics in 7 DaysLearn by IITians, NITians, Medical and Academic Advisor for each ExpertsDedicated studentDetailed Evaluationview performance all courses AdiChemistry IIT JEE atomic fluorine and neon rays in Angstrom units are respectively given by: It's (IIT JEE 1987) a) 0.72, 1.60Å, b) 1.60, 1.60Ã, c) 0.72, 0.72Ã, d) None of these logic and solution: the reported rays of noble gas elements are "van der doals rays ", which are 40% more than the actual ray atomic. Thus the Neon atomic radius must be much more fluorine. Note: It is not possible to obtain covalentes and metal radios for noble gas as they do not form links. Conclusion: in fact above is reflected in option "A". 2) Which of the following is the smallest in terms of size ã, (IIT JEE 1988) Å, 1) N3-Å, 2) O2- 3) F- 4) NA + Logic & solution:? An atom nof na atomic number, z (o) # protons 7 8 9 11 # electrons 7 8 9 11 Ã, ã, adda 3 electronics, | | | Å adda 2 electronsÅ | | | Å, ã, adda 1 electron | | | Å, Ã, subtract, 1 Electrona | | | One ion n3- o2- fanner + # electrons 10 10 10 10 z / # electrons Ã £ 7/10 = 0.7 ã, 8/10 = 0.8 ã, 9/10 = 0.8a ã, 11 / 10 = 1.1a The relationship between no. of protons (z) the number of electrons correlate effective nuclear attraction. The greater the value, the greater the attraction and smaller the size. For example, in na + ion, there is 1.1 proton for each electron. So the attraction is maximum and the ion is smaller than the species indicated. However, in ions n3-, there is only 0.7 protons for each electron. So the attraction is minimum and ion is great. Generalization: In the ISO-electronic species, the atomic size decreases with increasing the atomic number. Conclusion: The smallest is Na + .Ã, 3) The correct order of the rays is: Ã, (IIT JEE 2000) Ã, a) n